**Electro chemistry**

* Conductor : A substance which can allow the flow of electric current is called conductor .

 Ex: all metals , alloys , etc…

* Insulator: A substance which can not allow the flow of electric current is called insulator.

 Ex: wood, Paper , Plastic ,etc..

* Types of conductors: They are 2 types
1. Metallic Conductors: A conductor in which the electrons are carrying the current is called metallic conductors

 Ex: All metals, alloys ,etc…

1. Electrolytic Conductors: A conductor in which the ions are carrying the current is called electrolytic conductors .

 Ex: all acids, bases, salts , etc…

* Electrolyte: A substance which can decomposes into ions by the passage of electric current in molten state (or) aqueous solution is called electrolyte.

Ex: all acids , bases, salts , etc..

* Non-electrolyte: A substance which cannot decomposes into ions by the passage of electric current in molten state (or) aqueous solution is called non-electrolyte.

Ex: glucose , sugar , urea , benzene , etc…

Types of electrolyte:

* They are two types
1. Strong Electrolyte: A substance which can decomposes fastly into ions by the passage of electric current in molten state (or) aqueous solution is called strong electrolyte.

Ex: all strong acids , bases, salts , etc..

1. Weak Electrolyte: A substance which can decomposes slowly into ions by the passage of electric current in molten state (or) aqueous solution is called strong electrolyte.

Ex: all weak acids , bases, salts , etc..

Distinguish between metallic conduction and electrolytic conduction:

|  |  |
| --- | --- |
| **Metallic conduction**  | **Electrolytic conduction**  |
| 1. In this electrons are carrying current  | 1. In this the ions are carrying current  |
| 2. Ex: all metals , alloys , etc..  | 2. Ex: all acids , bases , salts , etc.. |
| 3.In this , the electrons moves from negative end to positive end | 3. In this , the ions moves towards the oppositely charged electrodes  |
| 4. In this , no chemical reactions takes place  | 4.In this , chemical reactions takes place |
| 5. In this , no flow of matter | 5. In this flow of matter. |

Electrolysis of fused(molten) NaCl:

* Electrolysis: the process of decomposition of an electrolyte into ions by the passage of electricity is called electrolysis.

Ex: electrolysis of fused NaCl



1. Anode: platinum rod
2. Cathode: platinum rod
3. Electrolyte: fused NaCl
4. Reactions:
5. Ionisation: NaCl ⇌ Na+ + Cl-
6. At anode: 2Cl- 🡪 Cl2 (↑) + 2e- (oxidation)
7. At cathode: 2Na+ + 2e- 🡪 2Na (reduction)

Note: The chemical equivalent of an electrode

 E = At.wt(M) / valency(n)



FARADAY’S LAW OF ELECTROLYSIS:

* Faraday’s 1st law : “The weight of the substance deposited at electrode is directly proportional to the quantity of electricity passed through the electrolyte” is called faraday’s 1st law .

i.e, W ∝ Q

 W ∝ (ct)

 🡪 w = ect

Where, w = weight of the substance deposited at electrode

 e= electro chemical equivalent

 c = current in amperes

 t = time in seconds

* Electro chemical equivalent (e) : The weight of substance deposited at electrode by passing 1 coulomb of electricity per 1 second is called electro chemical equivalent .

Faraday’s 2st law : “when the same quantity of electricity is passed through two (or) more electrolytes then the weight of the substances deposited at electrodes are directly proportional to their chemical equivalents” is called faraday's second law.

* i.e, W1 / E1 = W2 /E2.
* where, WI = weight of the 1st element

 EI = chemical equivalent of 1st element

 W2= weight of the second element

 E2 = chemical equivalent of 2nd element

Faraday:

* The quantity of electric current carried by 1mole of electrons (i.e, 6.023 x 1023 ) is called faraday.

 F= Ne

F= 6.023 x 1023 x 1.602 x 10-19

F =96,500 coulombs

RELATION BETWEEN ELECTRO CHEMICAL EQUIVALENT (e) AND CHEMICAL EQUIVALENT (E):

* Electro chemical equivalent (e) : The weight of substance deposited at electrode by passing 1 coulomb of electricity per 1 second is called electro chemical equivalent.
* e = M/n.F 🡪1
* Chemical equivalent (E): The weight of substance deposited at electrode by passing 1 Faraday (F=96,500 coulombs ) of electricity per 1 second is called chemical equivalent.

🡪E = M /n 🡪2

* From eq1

 e = M/n.F

 = (M/n)(1/F)

 = E . (1/F)

 🡺e = E/F

Formulae:

1. W = MCt/nF
2. C=WnF/Mt
3. t=WnF/MC
4. Faraday’s 1st law : W=eCt
5. Faraday’s 2nd law : W1 / E1 = W2 /E2.

Problems:

* Model-1
1. Calculate the weight of copper deposited when 10amperes of current is passed through a solution of CuSO4 for 10min?

Solution:

 Given that,W=?

 Current(C) =10amp

 Time (t)=10min=10x60=600sec

 Given that, CuSO4

 Atomic weight(M)=63.5

 Valency(n)=2

 Faraday(F)=96500Coulombs

 Since

 W=MCt/nF

 =(63.5x10x600)/(2x96500)

 =1.974gm

🡪The Weight of Copper Deposited(W)=1.974gm

* Model-2
	1. Calculate the current required in amperes to deposit 0.01gm of Zn in 1min 40 sec?

Solution:

 Given that, C=?

 Weight(W)=0.01gm

 Time(t)=1min 40sec

 =60+40=100sec

 Given that,Zn

 Atomic weight(M)=65

 Valency(n)=2

 Faraday(F)=96500coulombs

 Since

 C=WnF/Mt

 =(0.01x2x96500)/(65x100)

 =0.296amp

🡪 The required current is 0.296amp

* Model-3
1. How long would it taken to deposit 18gm of Al from an electrolytic cell containing Al2O3 using a current of 20amp?

Solution:

 Given that , t=?

 Weight(W)=18gm

 Current (C)=20amp

 Given that Al,

 atomic weight (M)=27

 Valency(n)=3

 Faraday(F)=96500Coulombs

 Since

 t=WnF/MC

 =(18x3x96500)/(27x20)

 =9650sec

🡪The required time =9650sec

* Model-4
1. A current of 2amp passing through AgNO3 Solution per 10min,deposits 1.4292gm of silver. What is the electro chemical equivalent of silver?

Solution:

 Given that,e=?

 weight(W)=1.4292gm

 current(C)=2amp

 time(t)=10min=10x60=600sec

 Since Faraday’s 1st law

 W=eCt

 e=W/Ct

 =1.4292/(2x600)

 =1.191x10-3gm/coulombs

🡪 e =1.191x10-3gm/coulombs

* Model-5
1. Calculate the weight of copped deposited from CuSO4  Solution , of 5.4gm of silver deposited from AgNO3 Solution by the same current ?[A.wt of Cu=63.5,At.wt of silver=108]

solutions:

 Given that: CuSO4

 The weight of copper(W1)=?

Chemical equivalent of copper(E1)=M1/n1

=63.5/2

 =31.75gm

Given that , AgNO**3**

 The weight of silver(W2)=5.4gm

Chemical equivalent of silver(E2)=M2/n2

 = 108/1

 = 108gm

Since , faraday’s second law

 W1/E1 = W2/E2

 W1=(W2XE1)/E2

 = (5.4X31.75)/108

 = 1.5875gm

 🡪The wt.of copper (W) = 1.5875gm

Galvanic cell (or) voltaic cell:

* A cell in which the chemical energy is converted to electric energy is called galvanic cell.
* Construction :



Explanation :

1. The galvanic cell contains two u-shaped beakers.
2. The first beaker containing Zn rod inserted into ZnSO4 solution which acts as anode.
3. The second beaker containing Cu rod inserted into CuSO4 solution which acts as Cathode.
4. The anode & cathode connected through a voltmeter
5. The ZnSO4 and CuSO4 solutions are seperated by a salt bridge.
6. The salt bridge is a u-shaped tube containing agar-agar gel with saturated KCl
7. The salt bridge prevents the physical mixing of two solutions.

Working:

1. At anode : Zn 🡪 Zn2+ + 2e-(oxidation)
2. At cathode: Cu2+ +2e-🡪 Cu(reduction)
3. Net reaction: Zn + Cu2+ 🡪 Zn2+ + Cu



* Role of salt bridge:
1. It prevents physical mixing of two Solutions
2. It prevents the accumulation of changes around the electrode
3. It allows the flow of current by completing the circuit
4. It maintains the electrical neutrality of the solution.

Distinguish between electrolytic cell & galvanic cell:



* Electrode Potential: The tendency of an electrode to loss (or) gain electrons is called electrode potential.
* These are two types.
1. Oxidation potential: The tendency of an electrode to Loss electrons is called oxidation potential.
2. Reduction potential: The tendency of an electrode to gain electrons is called reduction potential.
* STANDARD ELECTRODE POTENTIAL [Eo]: The tendency of an electrode of loss(or)gain electrons at 25°c with 1 molar concentration solution is called standard electrode potential .
* They are two types:
1. STANDARD OXIDATION POTENTIAL[SOP]: The tendency of an electrode to loss electrons at 25oC with 1 molar concentration solution is called standard oxidation potential [SOP]
2. STANDARD REDUCTION POTENTIAL[SRP]:The tendency of an electrode to gain electrons at 25oC with 1 molar concentration solution is called standard reduction potential [SRP]
* THE RELATION BETWEEN (SOP)&(SRP):

🡪 SOP = -SRP

ELECTRO CHEMICAL SERIES:

* The arrangement of metals in the increasing order of their standard reduction potential (SRP) is called electro chemical series.
* Significance:
1. A metal which is top in electro chemical series can acts as reducing agents
2. A metal which is lower in electro chemical series can acts as oxidising agents
3. It can be used to understand the reactivity of metals
4. It can be used to calculate the emf of galvanic cell.
5. It can be used for metallurgical operations
6. A metal which is top in electro chemical series can reacts with dilute acid to liberate H2 gas

Ex: Zn + 2HCl 🡪 ZnCl2 + H2 (↑)

* Electro motive force (emf): The potential difference between two electrodes of a galvanic cell is called emf
* It is denoted by “Eocell”
* emf(Eocell )=Eocathode - Eoanode

(or)

 = Eoright – Eoleft

* Units: volts

Importance:

1. If emf of the cell is positive then the cell can generate electricity
2. If emf of the is negative then cell cannot generate electricity

Problems:

* The standard reduction potential of “Zn” and “Cu” are -0.76 volts & +0.34volts respectively find the emf of cell (hint: [Zn/ Zn2+ (1M)// Cu2+ (1M)/Cu])

Solution: emf=?

 given that , “Zn” is anode

 “Cu” is cathode

 emf(Eocell )=Eocathode - Eoanode

 = 0.34 – (-0.76)

 = 0.34 + 0.76

 = 1.10volts

🡪 Emf =1.1volts

Practice problems:

* Calculate the weight of zinc deposited when 4amperes of current is passed through a solution of ZnSO4 for 30min?
* A current of 3amp passing through AgNO3 Solution per 20min,deposits 4gm of silver. What is the electro chemical equivalent of silver?
* The standard reduction potential of “Zn” and “Cu” are -0.76 volts & -0.40volts respectively find the emf of cell. Does the cell function? (hint: [Zn/ Zn2+ (1M)// Cu2+ (1M)/Cu])